

Hydrogen production via catalytic water splitting

Prospects of reducing greenhouse emission by hydrogen powered energy technologies





Mass energy densities for various fuels

Fuel	Hydrogen weight fraction	Ambient state	Mass Energy Density (MJ/kg)	
Hydrogen	1	Gas	120	
Methane	0.25	Gas	50	
Ethane	0.2	Gas	47.5	
Propane	0.18	Gas (Liquid) ¹	46.4	
Gasoline	0.16	Liquid	44.4	
Ethanol	0.13	Liquid	26.8	
Methanol	0.12	Liquid	19.9	

(1): A gas at room temperature, but normally stored as a liquid at moderate pressure

Facts about Liquid Fuels



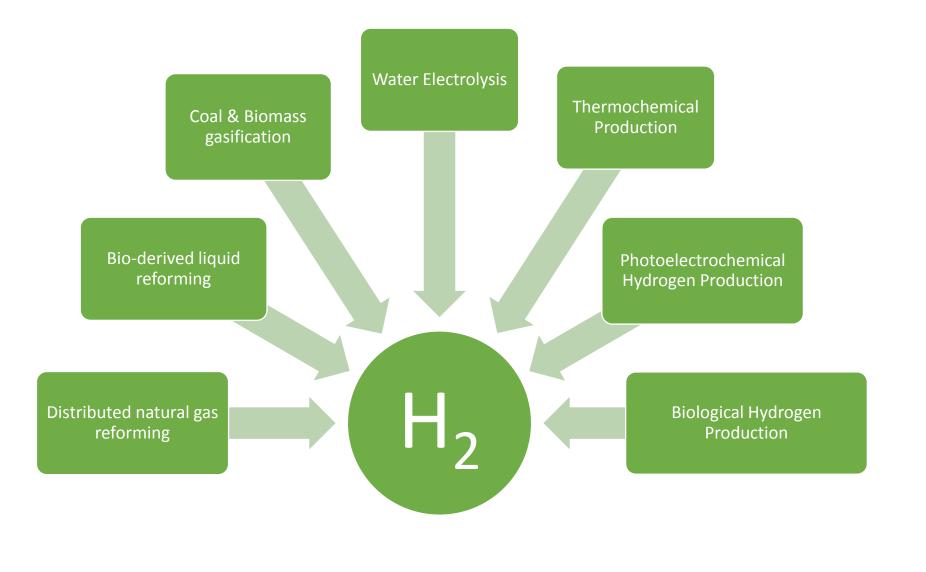
- For long-distance trucks and other heavy vehicles, as well as aviation, there is no good alternative to liquid fuels.
- Easy to transport and handle
- Examples are Gasoline, Diesel, kerosene and Alcohols

Facts about Hydrogen (H₂)



- Hydrogen is not an energy source. It is an energy carrier
- Hydrogen is a form of energy storage media
- We need primary energy sources such as sunlight, coal, natural gas or uranium to generate the power to extract Hydrogen from source materials such as natural gas and water.

Major Hydrogen Production Technology pathways









Major route of synthesis

Steam reforming of natural gas

Methane steam reforming: $CH_4 + H_2O \rightarrow CO + 3H_2$

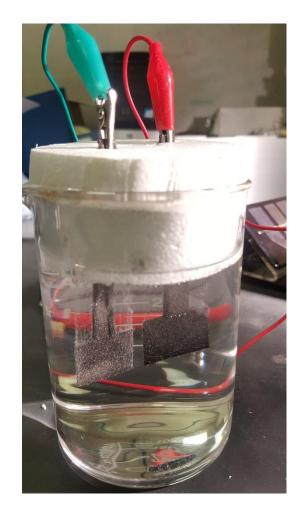
Water-gas-shift:

 $CO + H_2O \rightarrow CO_2 + H_2$





Water splitting is the cleanest way to produce Hydrogen



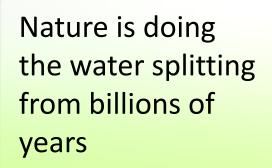
The Nature's secret: Turning sunlight into storable fuels

PHOTOSYNTHESIS



Photosynthesis

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The solution of your current problems might be hidden in the nature's secrets

+ 2H2

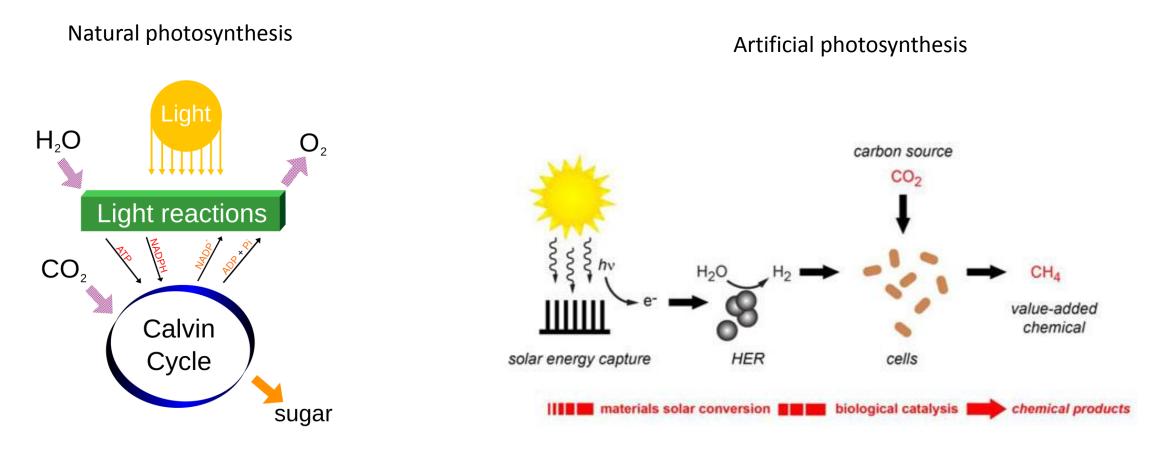
2(CH2O

0

$2H_2 + 2CO_2$

2H₂O

From natural photosynthesis to artificial photosynthesis

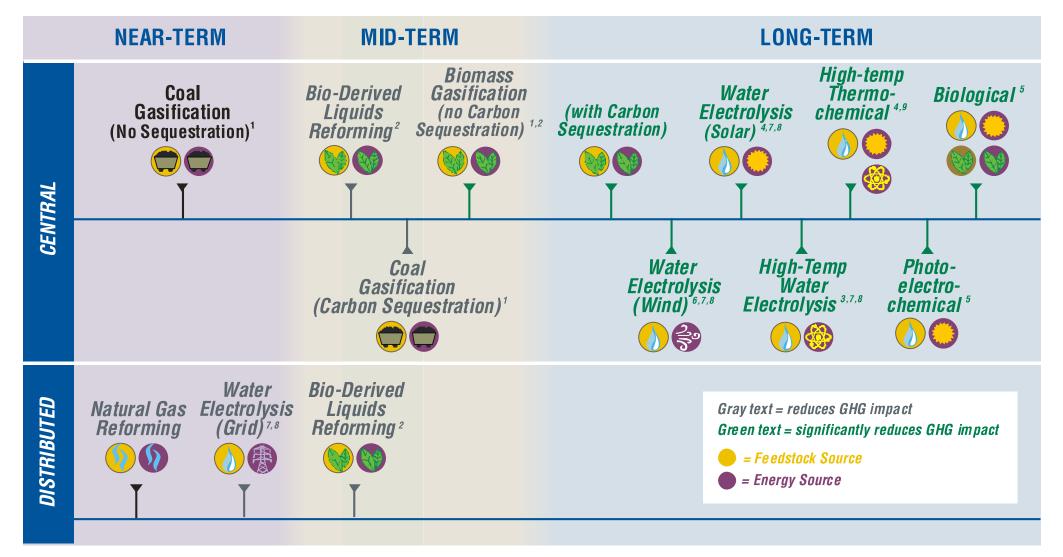


Natural Photosynthesis changes sunlight into chemical energy, splits water to liberate O_2 , and fixes CO_2 into sugar

Hybrid bioinorganic approach to solar-to-chemical conversion

Reference: E. M. Nichols, J. J. Gallagher, C. Liu, Y. Su, J. Resasco, Y. Yu, Y. Sun, P. Yang, M. C. Y. Chang and C. J. Chang, *Proceedings of the National Academy of Sciences*, 2015, **112**, 11461-11466.

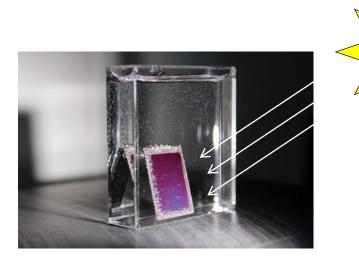
Technology Pathway Development Timelines, Feedstocks, and Energy Sources for Hydrogen Production



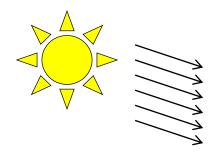
Enabling technologies under development by

- ¹ The Office of Fossil Energy
- 2 The Biomass Program
- 3 The Nuclear Hydrogen Initiative
- 4 The Solar Energy Technologies Program
- 5 The Office of Basic Energy Sciences
- ⁶ The Wind Program
- 7 The Geothermal Technologies Program
- ⁸ The Hydrogen Utility Group
- ⁹ The International Partnership for a Hydrogen Economy

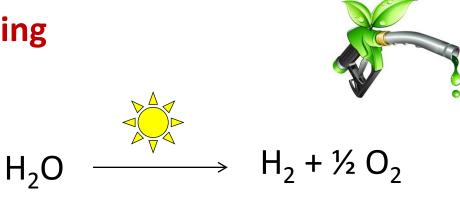
Hydrogen Production by catalytic water splitting



Water Splitting by sun light







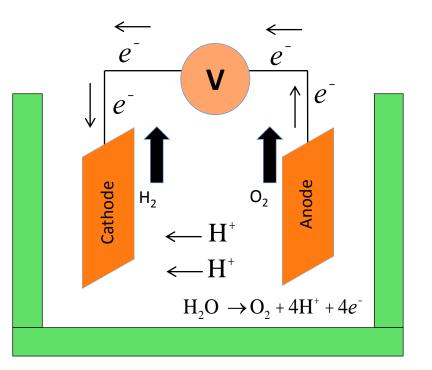
Features

- Hydrogen is a green fuel
- Hydrogen can be produced by water splitting
- The current human demand of electricity is 14 TW
- We need 16 TW by year 2050
- With an Olympic size pool of water, we can produce 43 TW electricity

How Hydrogen Generator works



1.23 V vs RHE at pH 0

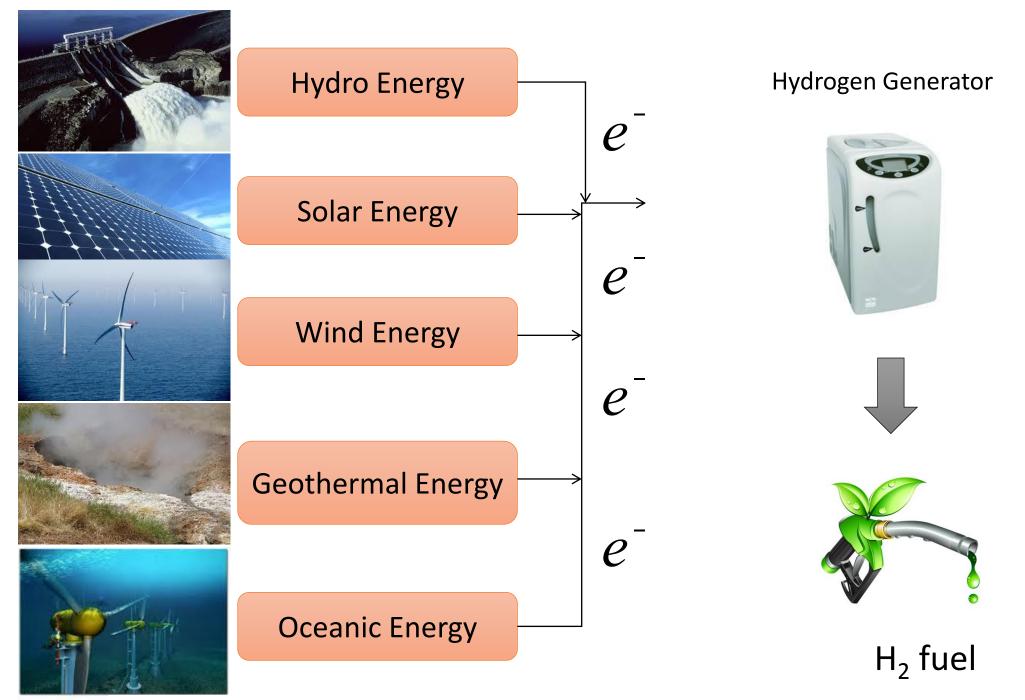


Electrochemical water splitting By using electricity



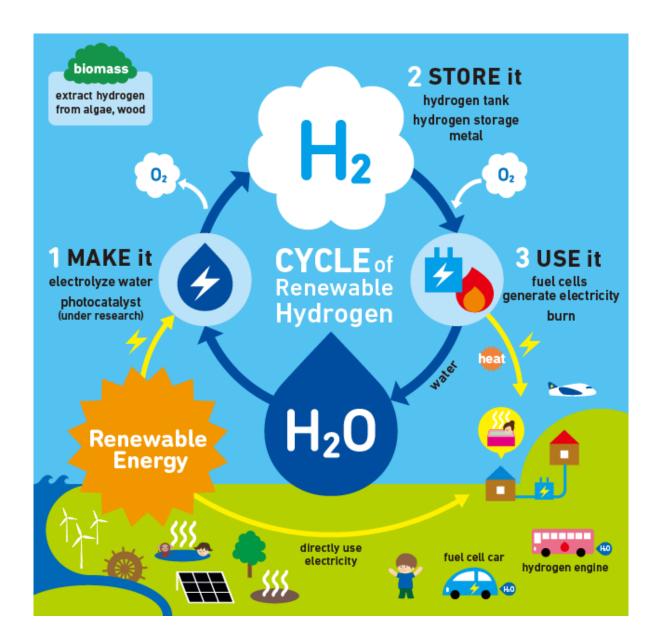


At Anode	$2\mathrm{H}_{2}\mathrm{O} \rightarrow \mathrm{O}_{2} + 4\mathrm{H}^{+} + 4e^{-}$		
At Cathode	$4\mathrm{H}^{+} + 4e^{-} \rightarrow 2\mathrm{H}_{2}$		
Overall	$2H_2O \Leftrightarrow O_2 + 2H_2$		



Ocean current turbine technolog energy generation on the OCS. In

Cycle of renewable hydrogen



Sustainable primary sources of Energy

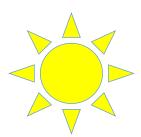
1- Wind Energy (can produce energy 24 hours a day)

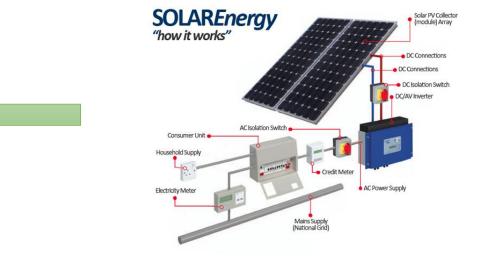
2- Solar Energy (can only produce energy from sunlight only)

Important Facts

- Panasonic has announced a rooftop solar panel, which is able to convert 22.5% of sunlight into electric energy.
- A kilogram of H₂—the energy equivalent of 4 liters of gasoline

Solar assisted water electrolysis Renewable Hydrogen



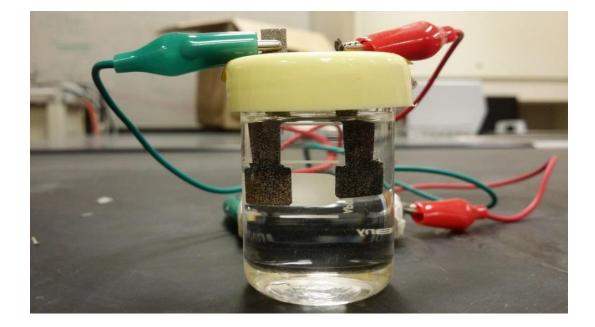


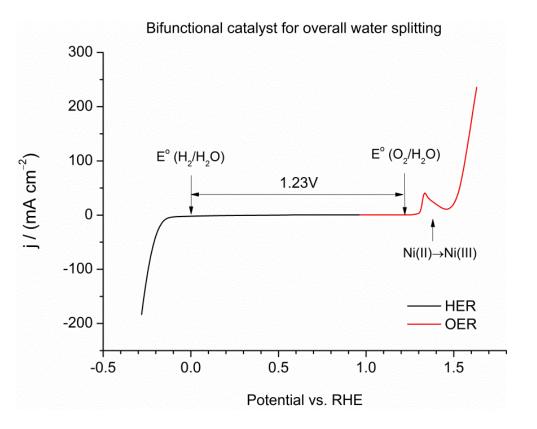
The three sizes of 900L Ovonics® Metal Hydride Hydrogen storage tanks



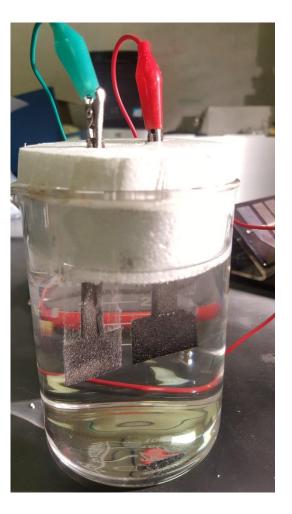
The hydrogen generators of the HG series (HG 30 | 60)

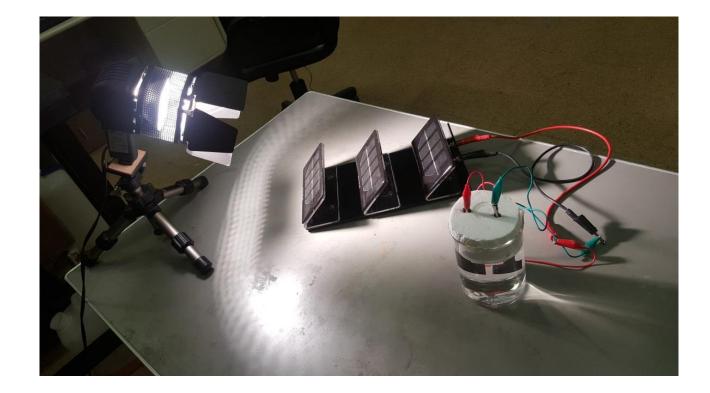
Catalytic water Splitting with 1.5V pencil cell





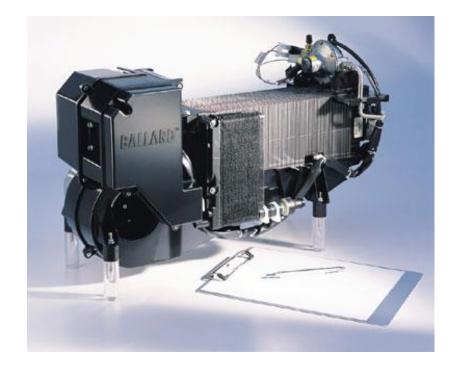
Solar assisted catalytic water splitting



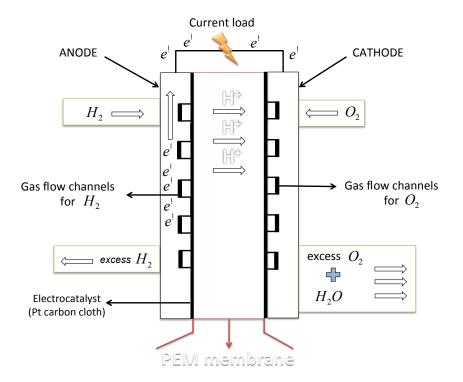


Hydrogen for power generation

Hydrogen powered PEMFC



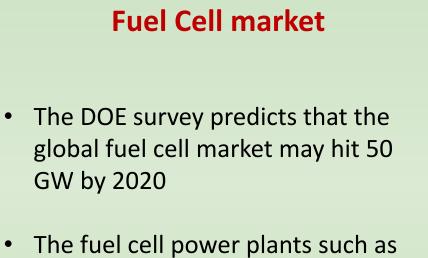
Polymer electrolyte Membrane Fuel Cell (PEMFC)



Schematic diagram of PEMFC system¹

1. Rana, U. A., M. Forsyth, et al. Towards protic ionic liquid and organic ionic plastic crystal electrolytes for fuel cells, Electrochimica Acta, 2012, 84, 213-222

Fuel Cell Stationary Power Plants



DFC 300 can be integrated to provide up to 900 kW



300 kW DFC 300 Fuel Cell power plant

Hydrogen powered Fuel Cell application in vehicle technology



General Motors Fuel Cell Hybrid vehicle (FCHV)



Hino and Toyota together launched Fuel Cell Hybrid vehicle (FCHV) bus

Hydrogen for our homes



Hydrogen in our homes

- A big question from the consumers of natural gas however is: would we want hydrogen mixed into the natural gas pipelines and would it be safe?
- In 1969, the coal gas containing up to 50% hydrogen was piped across the UK for lighting in homes, factories and for streetlights.
- The gas industry traces its origins back to 1792, when William Murdock used coal gas to light his home.
- The world's first gas company was formed in 1812 by Royal Charter, under the seal of King George III, and called the Gas Light and Coke Company. The company provided gas from coal gasification which contained up to 50% hydrogen, with the remainder being mostly carbon dioxide and carbon monoxide.
- The gas from coal gasification was used for lighting until the invention of the incandescent gas mantle in 1887.
- The first gas fire was produced in 1856, meaning heat could also be produced from burning the hydrogen in coal gas.

Reference: http://www.cleantechinvestor.com/portal/fuel-cells/11024-hydrogen-in-our-homes.html

Hydrogen in our homes continues.....

- The eventual conversion of the UK gas infrastructure to methane (CH₄) began in 1969 and was completed five years later by 1974.
- This led to the consumer setup we have today, where natural gas is used for heating, hot water provision and cooking, but it also allowed for electricity production using gas turbines.
- The methane based gas generators have high levels of carbon footprint.
- Hardy^{*} also emphasized that in order for the UK government to meet its emissions reduction targets of 80% by 2050, reverting back to a piped hydrogen network would an ideal way to proceed.

*Peter Hardy, Technical Services Manager at the Institute of Gas Engineers & Managers

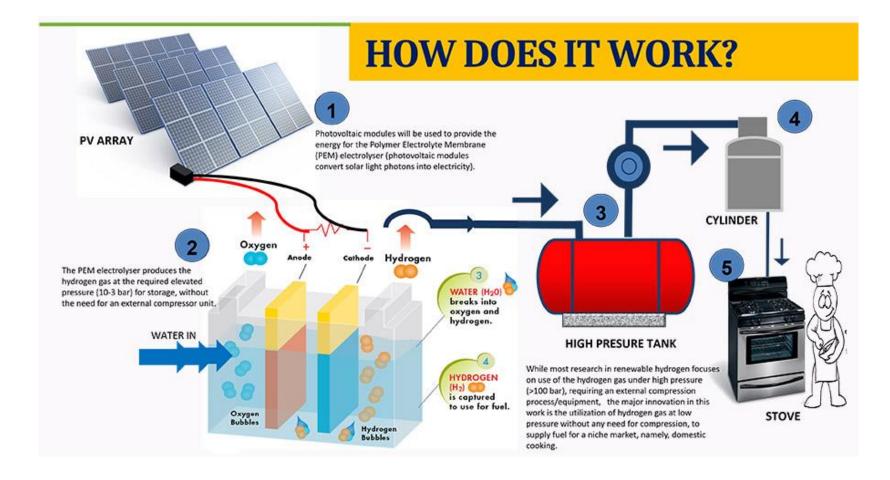
Reference: http://www.cleantechinvestor.com/portal/fuel-cells/11024-hydrogen-in-our-homes.html

Simultaneous Reduction of CO₂ and Splitting of H₂O

Syngas to SynFuel:

Syngas or Synthetic gas is a gas mixture consisting primarily of hydrogen (H_2) and carbon monoxide (CO).

Sustainable Hydrogen powered Stove Project (Jamaican Govt. initiative¹)

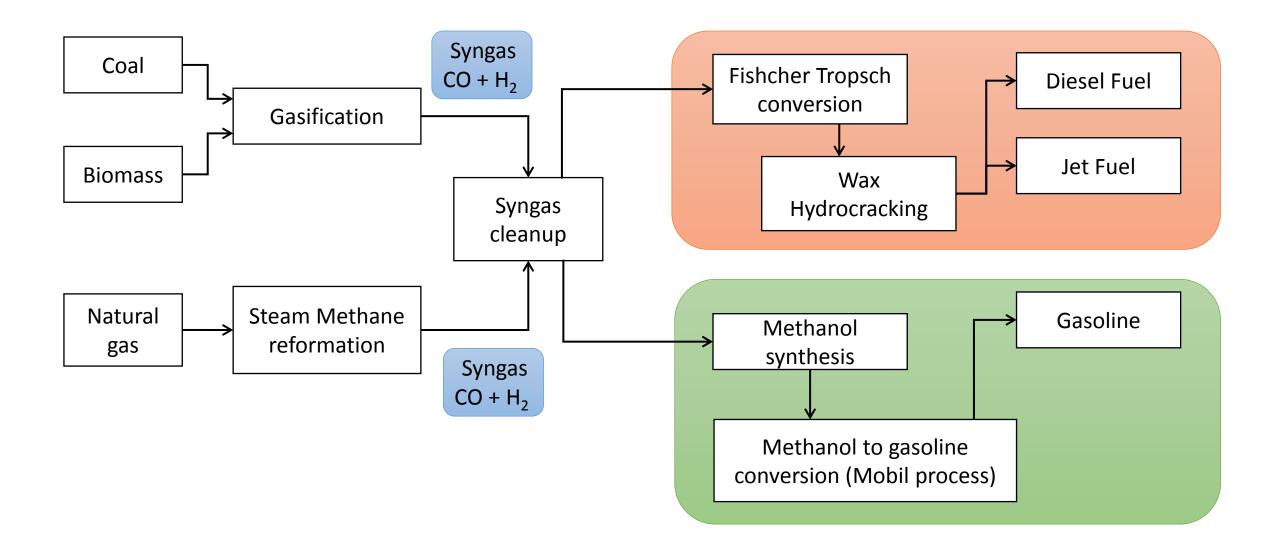


Drawbacks

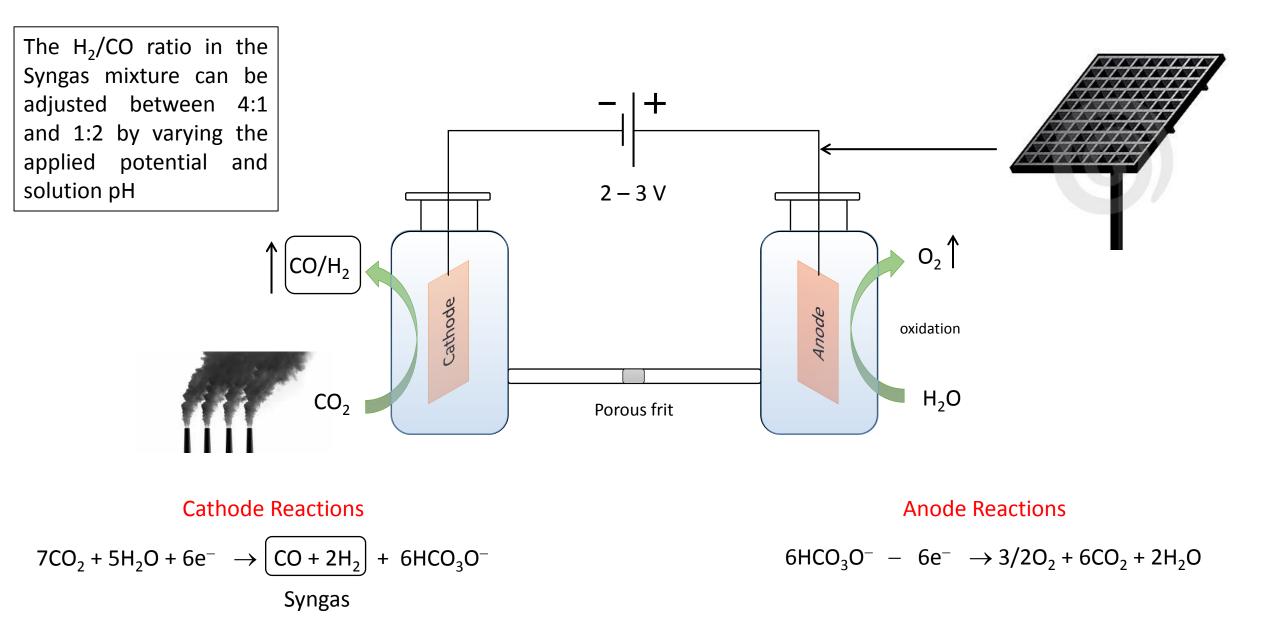
To produce hydrogen gas for cooking purposes will require large PEM electrolyzers, which will contribute to significantly high cost of the complete system.

1. http://www.solarhydrogen.utechsapna.com/Home/About

Synthetic Fuels manufacturing process



Simultaneous Reduction of CO₂ and Splitting of H₂O



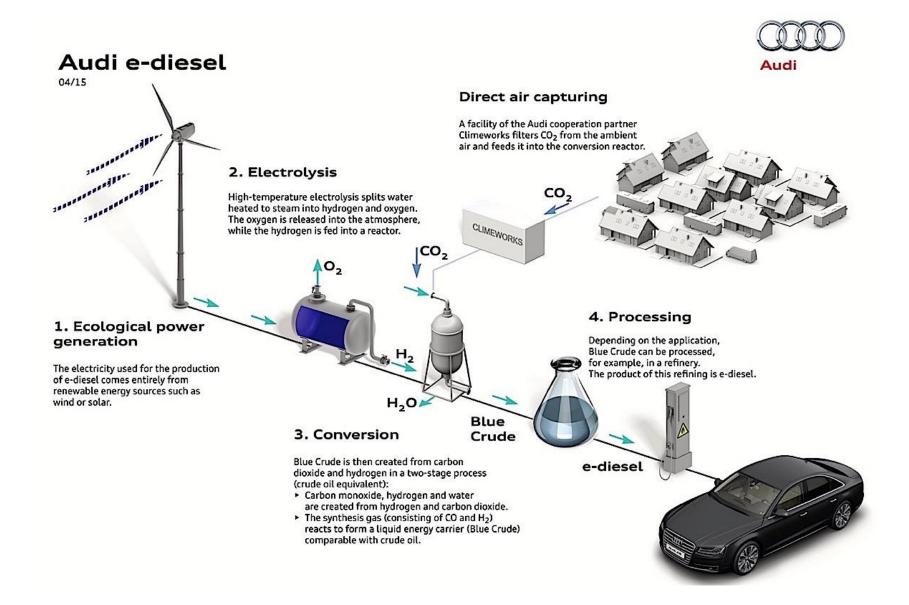
There's too much carbon dioxide in the air. Why not turn it back into fuel?



Carbon footprint

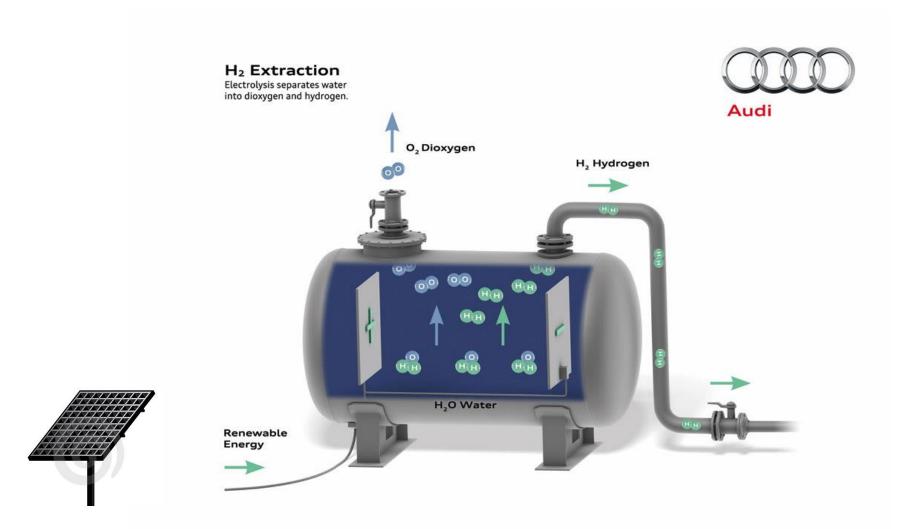
Reference: http://www.sciencemag.org/news/2015/09/feature-there-s-too-much-carbon-dioxide-air-why-not-turn-it-back-fuel

Carbon neutral fuels: Synthetic Diesel from H₂O and CO₂

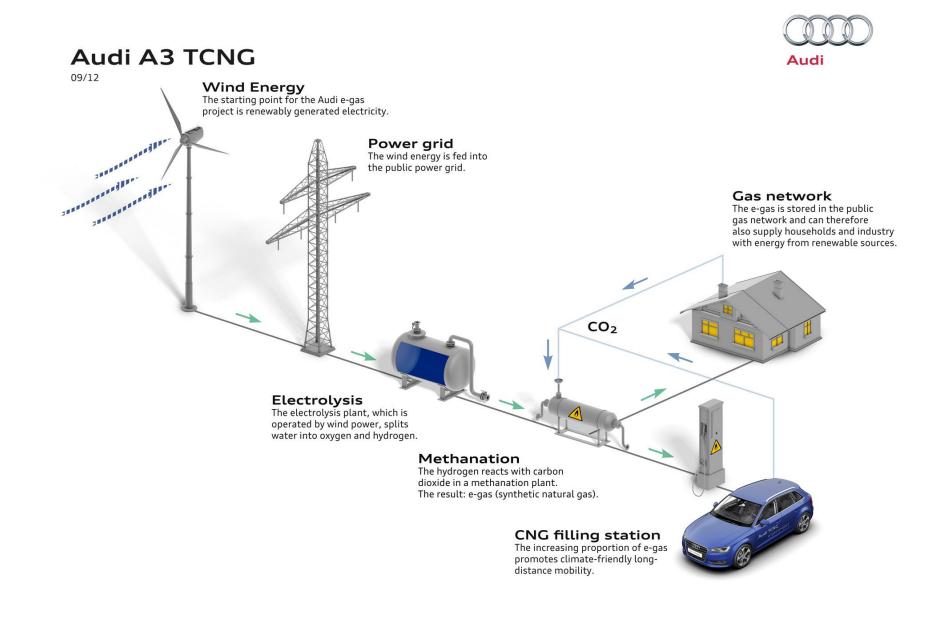


Water Electrolysis

The key to renewable hydrogen for synthetic diesel and synthetic fuel gas (methane)

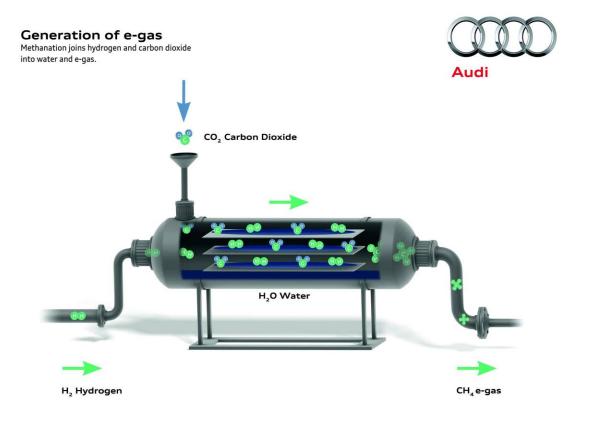


Carbon neutral fuels: Synthetic Methane (CH₄) from H₂ and CO₂



Audi Transforming Wind Power into eGas (synthetic methane) for New A3 gTron

- Synthetic methane can be produced from the methanation (reacting CO₂ with H₂)





Questions

In acid and neut	tral electrolytes
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In alkaline electrolytes

Overall	$2H_2O \Leftrightarrow O_2 + 2H_2$	E ⁰ = -1.23 V <i>vs.</i> RHE	Over	$all 2H_2O \Leftrightarrow O_2 + 2H_2$	E ⁰ = -1.23 V <i>vs.</i> RHE
At Anode	$2H_2O \rightarrow O_2 + 4H^+ + 4e^-$	E ⁰ = 1.23 V <i>vs.</i> RHE	At Anode	$4\mathrm{OH}^{-1} \rightarrow \mathrm{O}_2 + 2\mathrm{H}_2\mathrm{O} + 4e^-$	E ⁰ = 0.40 V <i>vs</i> . RHE
At Cathode	$4\mathrm{H}^{+} + 4e^{-} \rightarrow 2\mathrm{H}_{2}$	E ⁰ = 0.00 V <i>vs.</i> RHE	At Cathode	$4\mathrm{H}_{2}\mathrm{O} + 4e^{-} \rightarrow 2\mathrm{H}_{2} + 4\mathrm{OH}^{-1}$	E ⁰ = -0.83 V <i>vs.</i> RHE